



Nano-engineering of three-dimensional core/shell nanotube arrays for high performance supercapacitors



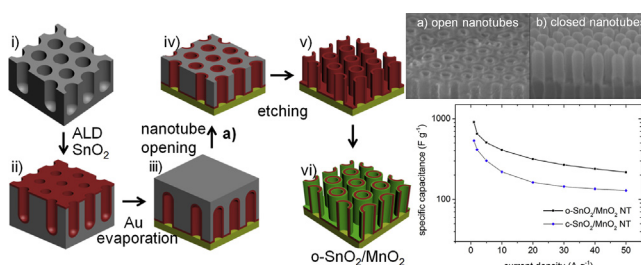
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HIGHLIGHTS

- Free-standing three-dimensional $\text{SnO}_2/\text{MnO}_2$ core/shell nanotubes.
- Demonstration of an innovative template-based fabrication technique with high structural controllability.
- Novel strategy to selectively fabricate open-end and closed-end nanotubes.
- Excellent electrochemical performance compared to other MnO_2 -based nanostructures.
- High specific capacitance 910 F g^{-1} at 1 A g^{-1} and a good rate capability.

GRAPHICAL ABSTRACT



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ABSTRACT

Large-scale arrays of core/shell nanostructures are highly desirable to enhance the performance of supercapacitors. Here we demonstrate an innovative template-based fabrication technique with high structural controllability, which is capable of synthesizing well-ordered three-dimensional arrays of $\text{SnO}_2/\text{MnO}_2$ core/shell nanotubes for electrochemical energy storage in supercapacitor applications. The SnO_2 core is fabricated by atomic layer deposition and provides a highly electrical conductive matrix. Subsequently a thin MnO_2 shell is coated by electrochemical deposition onto the SnO_2 core, which guarantees a short ion diffusion length within the shell. The core/shell structure shows an excellent electrochemical performance with a high specific capacitance of 910 F g^{-1} at 1 A g^{-1} and a good rate capability of remaining 217 F g^{-1} at 50 A g^{-1} . These results shall pave the way to realize aqueous based asymmetric supercapacitors with high specific power and high specific energy.

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1. Introduction

Energy storage and energy management are one of the key challenges of mankind in this century in order to meet the requirement of an increasing intermittent energy supply. The demand of energy storage systems include small storage systems for energy harvesting applications [1], medium size applications such

as energy storage systems for electrical mobility [2], and large scale systems leveling peaks in power grids [3]. In all these applications supercapacitors play a crucial role and are discussed as a desirable solution to address these emerging challenges. In the past decade various approaches have been attempted to increase the performance of supercapacitors. First, severe effort was made to increasing the electrode surface area by reducing the surface morphology into the nanometer regime and fabricating nanostructured electrodes [4–6]. And second, new electrode materials, such as activated carbons [7,8], and pseudocapacitive materials including conductive polymers [9–11] and metal oxides [4,12–15]

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were introduced. [6,16–18] So far two major challenges in the field of supercapacitors exist: (i) highly promising supercapacitor materials such as MnO_2 suffer from high internal resistance ($\sim 10^2$ – $10^6 \Omega \text{ cm}^{-1}$), which limits the electrochemical performance in supercapacitor devices; (ii) it is important to realize a high controllability of all structural parameters to optimize the electrochemical properties of supercapacitors, so as to identify the ideal configuration of high performance devices. In order to solve these challenges, novel functional nanostructures and material compositions (and related fabrication processes) need to be developed. An attractive solution is offered by preparing large-scale arrays of one-dimensional (1D) core/shell nanostructures including nanowires and nanotubes. Such nano-engineered materials combine the unique properties of two materials (i.e., core and shell). Usually, the highly conductive core provides a fast electron transport through the entire electrode and the thin shell ensures a well utilization of the energy storage material and a short ion diffusion. Template-based techniques are efficient methods to prepare large-arrays of 1D nanostructures [19–23], especially for core/shell structures [24,25]. However, the level of controllability of the template-based techniques is partially limited to a few parameters and largely depends on the inherent structural nature of the template itself. Atomic layer deposition (ALD), as an efficient process to prepare conformal films or tubular structures, provides a good solution to further adjust more structural parameters.

Here, we demonstrate an innovative technique to realize highly ordered free-standing three-dimensional (3D) arrays of $\text{SnO}_2/\text{MnO}_2$ core/shell nanotubes with precise controllability of many structural parameters. This includes not only diameter, spacing, and length of the core/shell nanostructures originated from the template fabrication process, but also the wall thickness of the core and shell nanotubes. Moreover, advanced core/shell nanostructures including open-end and closed-end nanotubes (denoted as o- for open-end and c- for closed-end nanotubes) are achieved, so as to investigate and optimize the electrochemical mechanism for supercapacitor applications. Very importantly, based on the property optimization of the 3D core/shell nanostructures, we obtained an electrochemical performance that is among the highest reported values for MnO_2 -based nanostructures.

2. Experimental section

2.1. Preparation

Porous anodic aluminum oxide nano-templates were fabricated using a two-step anodization process from 99.999% pure aluminum foils with a second anodization time of 5–15 min followed by a pore widening process in a 5 wt% H_3PO_4 solution for 15 min [26,27]. Such templates were used for the growth of the SnO_2 nanotube arrays by a PicoSun ALD system according to the following procedure. The reaction chamber was heated to 250 °C, and SnCl_4 and H_2O were selected as the precursors. First SnCl_4 precursor was pulsed for 1 s and purged for 4 s, followed by a 2 s pulse and 8 s purge of H_2O . This procedure was repeated for 500–1500 times depending on the desired wall thickness of SnO_2 nanotubes. After this process, a 100 nm thick layer of gold was evaporated onto the sample by physical vapor deposition to support the SnO_2 nanotube array. Afterward the remaining Al was removed by a saturated solution of CuCl_2 . In the next step, the template was removed. Hereby different processes for open-end and closed-end nanotube arrays were used. For the fabrication of closed-end nanotubes the entire templates was removed in a 5 wt% NaOH solution for 120 min. In order to successfully fabricate open-end SnO_2 nanotubes, the templates were first partially etched for 20 min to only liberate the nanotubes top parts. Then the top-ends of the nanotubes were removed by

scratching the surface with a sharp scalpel. Thereafter the sample was placed again in a 5 wt% NaOH solution to remove the remaining template. Finally the SnO_2 nanotube array was attached to a stainless steel holder.

The as-prepared open-end and closed-end SnO_2 nanotube arrays were coated by MnO_2 shells. A thin MnO_2 layer was formed by electrochemical deposition in an electrolyte containing 0.1 M $\text{Mn}(\text{Ac})_2$ and 0.1 M Na_2SO_4 , using 1 V versus a Ag/AgCl reference electrode. A Pt foil was used as a counter electrode. The charge passed was controlled during the deposition process within the range of 130–260 mC. Afterward the samples were cleaned with water and dried.

The entire fabrication process with all steps is shown schematically in the supplementary video.

Supplementary data related to this article can be found online at <http://dx.doi.org/10.1016/j.jpowsour.2013.12.029>.

2.2. Electrochemical and physical characterizations

All electrochemical measurements were performed in a three electrode cell with a 1 M Na_2SO_4 electrolyte, an Ag/AgCl reference electrode, and a Pt foils as the counter electrode. The CV and charge/discharge curves were measured from 0 to 0.9 V vs Ag/AgCl at scan rates of 2, 5, 10, 20, and 50 mV s^{-1} and current densities of 1, 2, 5, 10, 20, 30, 40, 50 A g^{-1} , respectively. The EIS was performed at 0.1 V with a 10 mV pulse in the frequency range of 100 kHz–0.05 Hz. All measurements were performed on a Bio-Logic VSP electrochemical work station.

The structural characterization was carried out using SEM (Hitachi S-4800) and TEM (Libra 200 FE).

X-ray photoelectron spectroscopy (XPS) was performed under an emission angle of 53° using a monochromated AlK α source and a hemispherical electron analyzer.

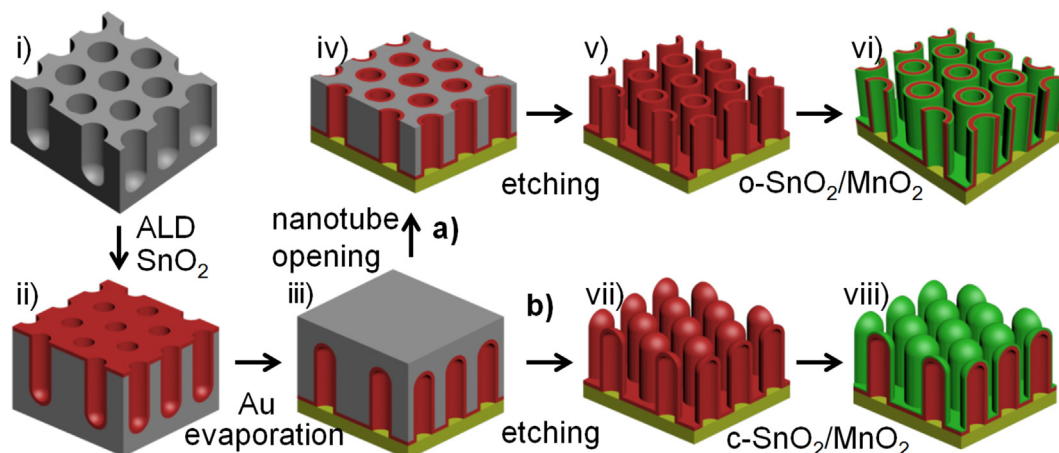
3. Results and discussion

3.1. Characterization of the $\text{SnO}_2/\text{MnO}_2$ core/shell nanotube array

The fabricated $\text{SnO}_2/\text{MnO}_2$ core/shell nanotube arrays are prepared by the ALD process of SnO_2 into a porous anodic aluminum oxide nano-template and a subsequent conformal coating by MnO_2 , as illustrated in Scheme 1. Thereby it is possible to selectively control the synthesis of o- $\text{SnO}_2/\text{MnO}_2$ (process a in Scheme 1) and c- $\text{SnO}_2/\text{MnO}_2$ core/shell nanotubes (process b in Scheme 1). This selectivity is a salient feature of the fabrication process and highly desirable for supercapacitor applications. The possibility of fabricating open-end nanotube arrays enables the access of the total inner and outer surface area of a nanotube for charge storage and thus is superior to closed-end nanotube arrays, which possess a lower accessible surface area.

The morphology of the prepared SnO_2 core nanotubes is analyzed by scanning electron microscopy (SEM) and transmission electron microscopy (TEM), as shown in Fig. 1. The SEM image in Fig. 1a clearly shows a highly ordered free-standing array of o- SnO_2 nanotubes (for c- SnO_2 nanotubes see Fig. S1). The outer diameter and the inter-tube distance of the nanotubes are about 60 nm and 105 nm respectively, which are inherited from the structural parameters of the templates. Different lengths of nanotubes between 600 nm and 1800 nm are obtained by using templates with different thicknesses, indicating that the length of nanotubes can be tuned by the template thickness.

The opening process of the SnO_2 nanotubes is shown in Fig. 2. First, the anodic aluminum oxide nano-template is partially etched to liberate the closed nanotube top-ends. Then the liberated closed nanotube top-ends are removed by scratching the samples surface



Scheme 1. Illustration of the fabrication processes of a) open-end o-SnO₂/MnO₂ nanotube arrays and b) closed-end c-SnO₂/MnO₂ nanotube arrays. i) Starting from porous anodic aluminum oxide nano-template, ii) conformal coating of template with SnO₂ by atomic layer deposition, iii) turning sample and fabrication of current collector by e-beam evaporation of gold; process a): iv) partial etching of porous anodic aluminum nano-template and nanotube opening by scratching process, v) removing of template to gain highly ordered free-standing SnO₂ nanotube array, vi) electrochemical coating with MnO₂; process b): vii) complete removing of template, viii) electrochemical coating with MnO₂.

with a sharp scalpel, resulting in open-end nanotubes (please refer to the supplementary video). The wall thickness and the inner diameter of the SnO₂ nanotubes are investigated by TEM measurements (Fig. 1c). The wall thickness of about 8 nm is evaluated from the contrast profile in the inset of Fig. 1c (using the similar method as described in Ref. [28]). It depends on the total cycle number during the ALD fabrication process and the growth rate, which is estimated to be approximately 0.08 Å/cycle. It is found that 8 nm is an optimal trade-off of wall thickness to maximize the overall inner surface area (*i.e.*, enlarge the inner-tube diameter) and in the same time to keep high mechanical stability of the SnO₂ nanotubes (*i.e.*, thicken the tube walls). Therefore, for the

nanotubes with about 60 nm outer diameter, 105 nm inter-tube distance and 8 nm tube thickness, the inner diameter and the inter-tube spacing are about 44 nm and 45 nm, respectively. Hence there are sufficient inner- and outer-tube spaces for the next step of conformal coating of MnO₂.

Subsequently the SnO₂ nanotubes are conformal coated by MnO₂ in an electrochemical deposition process. The formed MnO₂ shell layer is investigated by TEM and SEM measurements. Fig. 3a shows a TEM image of a single SnO₂ nanotube covered entirely by a thin layer (10–20 nm thickness) of MnO₂. The elemental distribution across the SnO₂/MnO₂ nanotube is investigated by EDX line scan and verifies the presents of Sn, Mn, and O. The elemental distribution of Sn shows two peaks located at the wall positions of the nanotube, whereas Mn is distributed more uniformly across the nanotube (this might originates from the fact that MnO₂ was coated on both the inner and outer surfaces of SnO₂ tubes). The SEM image in Fig. 3b shows that the SnO₂ nanotubes are covered completely by MnO₂. The difference between the MnO₂-coated nanotubes and uncoated bare SnO₂ nanotubes is presented clearly in Fig. 3b. The coated nanotubes possess much rougher surfaces comparing to uncoated tubes, which is a typical results of electrochemical deposited MnO₂. Well conformal coating of MnO₂ on SnO₂ nanotubes can only be achieved by using optimized fabrication

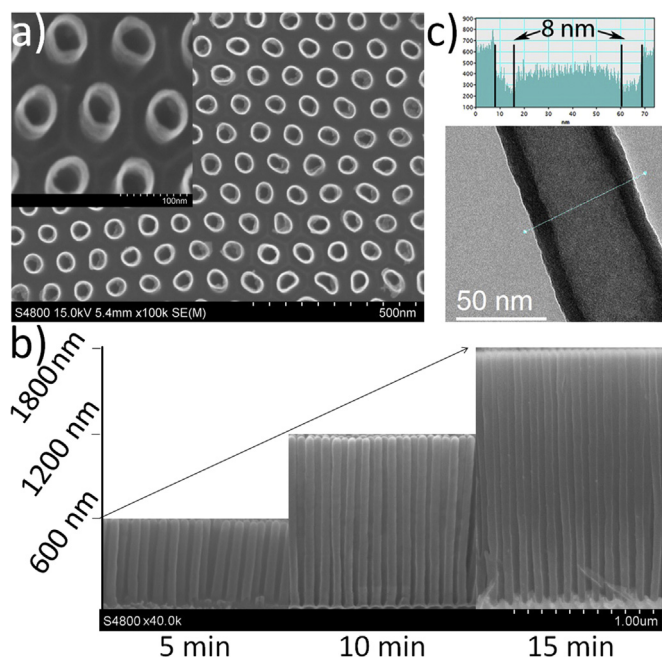


Fig. 1. a) Top view SEM image of an o-SnO₂ nanotube array. The inset shows the inter-tube distance $d_{\text{int}} = 105$ nm and inter-tube spacing $d_{\text{spac}} = 44$ nm b) cross-section SEM image showing the length of the nanotubes in dependence of the template anodization time (*i.e.*, thickness of the template). c) TEM image, showing the hollow nature of the nanotubes. The inset shows the contrast profile across the nanotube and reveals a wall thickness of around 8 nm.

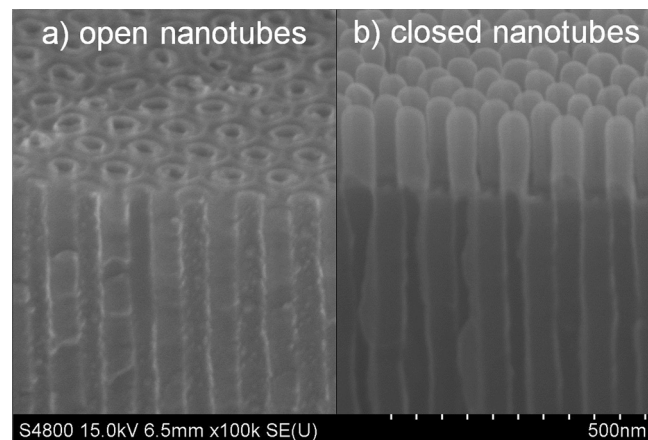


Fig. 2. a) Open nanotube array after removing the closed top-ends of the nanotubes by scratching the surface with a sharp scalpel; b) partially etched nanotube array showing the liberated closed top-ends before removing.

conditions and controlling the amount of material being deposited (an excessive deposition of MnO_2 is leading to an in-conformal coating as shown in Fig. S2). Finally an optimal window of 130–260 mC flown charge in electrochemical deposition is determined for well conformal coating of MnO_2 .

3.2. Electrochemical performance

The electrochemical performance of the prepared core/shell nanostructures is investigated in a three-electrode setup. The measuring cell includes a platinum foil as a counter electrode, an Ag/AgCl reference electrode and a water-based electrolyte containing 1.0 M Na_2SO_4 . Fig. 4a shows the cyclovoltammetry (CV) curves of o- $\text{SnO}_2/\text{MnO}_2$ nanotube arrays. All CV scans were performed in a potential range from 0 to 0.9 V vs Ag/AgCl at scan rates of 2, 5, 10, 20, and 50 mV s^{-1} and show scaling behavior of the current as the scan rate is increased. The shape of the CV curves reassembles the expected shape for pseudocapacitive charge storage in MnO_2 and indicates good capacitance behavior of our sample. During a standard cycle voltammetry testing, Mn undergoes redox reactions on the surface of samples (MnO_2 shell in our case). Thereby the valance state of Mn shall be reduced in the charging process and increased in the discharging process. This process is caused by the intercalation/deintercalation of protons (H^+) or alkali cations (Na^+) into/out of the MnO_2 shell. This reaction can be described as [29]:



Galvanostatic charge/discharge curves of as-prepared o- $\text{SnO}_2/\text{MnO}_2$ nanotube array are recorded at different current densities to evaluate the specific capacitance and rate capability, as shown in Fig. 4b. The specific capacitance is calculated for various current densities from the associated discharge curve according to the following equation:

$$c = I / \left(\frac{\Delta V}{\Delta t} \times m \right) \quad (3)$$

with c (F g^{-1}), I (A), $\Delta V/\Delta t$ (V s^{-1}), and m (g) being the specific capacitance, discharge current, gradient of discharge curve

obtained by dividing the width of the potential window (i.e., 0.9 V) by the discharge time, and mass of the active material, respectively. It is noteworthy that the gained specific capacitance of the core/shell structures can be attributed to the charge storage capability of the MnO_2 shell. [30] This assumption is justified because the SnO_2 nanotubes are conformally covered by MnO_2 during the shell formation and therefore can rarely contribute to the charge storing capability because the charges are stored near or at the surface of the electrode. [30] Moreover, in the Supplementary information we discussed the possible influence of the SnO_2 on the electrodes performance. The graph in Fig. 4d displays the obtained specific capacitance at different current densities and compares the rate capability of o- $\text{SnO}_2/\text{MnO}_2$ and c- $\text{SnO}_2/\text{MnO}_2$ core/shell nanotubes. It can be seen that the o- $\text{SnO}_2/\text{MnO}_2$ nanotube array exhibits an extremely high specific capacitance of 910 F g^{-1} at 1 A g^{-1} and remains a high specific capacitance of 217 F g^{-1} at 50 A g^{-1} . As shown in Fig. 5, this performance outnumbers the performance of pure nanostructured MnO_2 electrodes. And very importantly, it is among the highest reported values of MnO_2 -based core/shell nanostructures. Beyond that the prepared core/shell structure exhibits a high areal capacitance of up to 0.092 F cm^{-2} at 1 A g^{-1} and 0.022 F cm^{-2} at 50 A g^{-1} , which is higher than reported values of other core/shell structures [31]. In addition the reported structure obtains a high volumetric capacitance of 511 F cm^{-3} at 1 A g^{-1} and 122 F cm^{-3} at 50 A g^{-1} . The performance of the c- $\text{SnO}_2/\text{MnO}_2$ nanotube arrays remains behind the values achieved by o- $\text{SnO}_2/\text{MnO}_2$ nanotube arrays at same mass loadings, as presented in Figs. 4d and 5. It is believed that this difference could be attributed to the lower available surface area of c- SnO_2 electrodes and a longer ion diffusion length. It clearly shows that a highly porous open-end nanotube structure as a core material is superior to core materials based on closed-end nanotubes and nanowires.

The investigation of the long-time cycling stability of the o- $\text{SnO}_2/\text{MnO}_2$ nanotube array is shown in Fig. 4c. The electrode is consecutively charged and discharged for 2600 cycles at 5 A g^{-1} . The electrode capacitance decreases during the first 100 charge/discharge cycles. Thereafter the charge/discharge profile is constant for 2500 cycles, which is confirmed by the charge/discharge curves in the inset of Fig. 4c. This result demonstrates that the o- $\text{SnO}_2/\text{MnO}_2$ nanotube arrays are capable of accomplishing thousands of charge/discharge cycles and are ideal candidates for long-time applications.

The electrochemical performance of the o- $\text{SnO}_2/\text{MnO}_2$ nanotube array is further investigated by electrochemical impedance spectroscopy, as shown in the Nyquist plot in Fig. S4. The impedance plot is composed of a steep line in the low frequency region and a semi-circle in the high frequency region. The almost vertical line in the low frequency region reveals the good capacitive behavior of the electrode material due to fast and reversible reactions occurring at or near the surface of the electrode. The high frequency region gives evidence about the equivalent series resistance (R_s) at the low x-axis interception and about the charge transfer resistance (R_{ct}) at the interception at $(R_s + R_{ct})$ [32]. The equivalent series resistance and the charge transfer resistance of the o- $\text{SnO}_2/\text{MnO}_2$ nanotube array are measured to be 2.09Ω and 1.53Ω , respectively, and are lower or comparable to other reported core/shell structures [33,34]. In contrast, electrodes consisting of pure MnO_2 generally show higher resistance values due to the high internal resistance of MnO_2 . For comparison, the electrochemical impedance spectrum obtained from a thin MnO_2 layer on a gold substrate with same mass loading shows remarkably higher charge-transfer resistance compared to the core/shell structure, see Fig. S4. These results obtained from EIS indicate that the o- $\text{SnO}_2/\text{MnO}_2$ core/shell structure with its conductive core and thin shell is a favorable

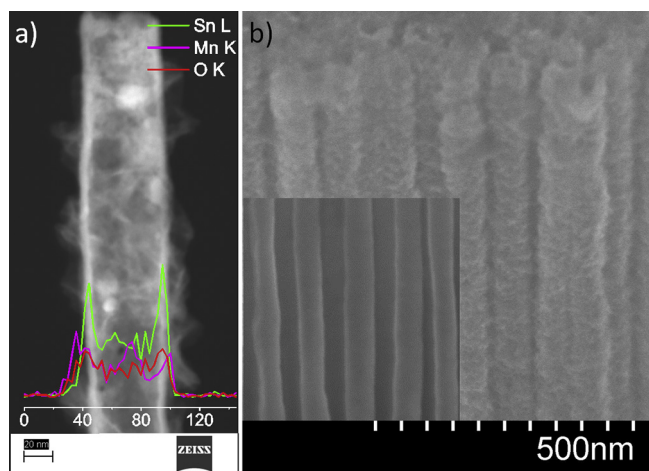


Fig. 3. Morphology of o- $\text{SnO}_2/\text{MnO}_2$ core/shell nanotube. a) Cross-section TEM image with EDX line scan showing the elemental distribution of Mn, Sn, and O across a single core/shell nanotube. b) SEM image of MnO_2 -coated SnO_2 nanotubes compared with bare SnO_2 nanotubes shown in the inset. Scale bar is valid for both images.

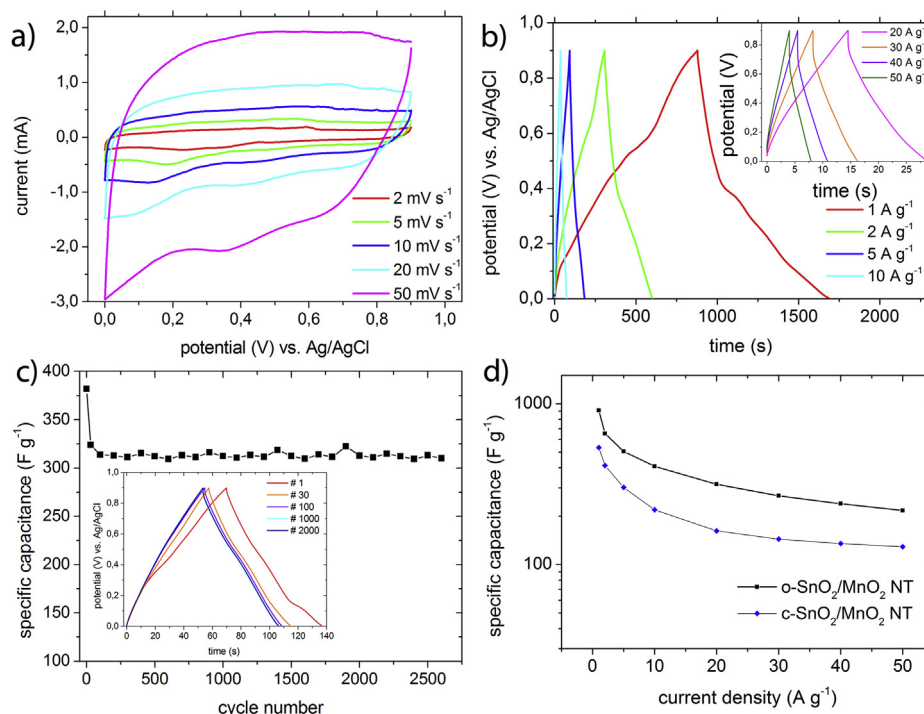


Fig. 4. Electrochemical performance of the o-SnO₂/MnO₂ core/shell nanotube array with a nanotube length of 1.8 μm . a) CV curves at different current densities, b) charge/discharge profiles at different current densities, c) long-time cycling stability at 5 A g⁻¹ including various charge/discharge curves in the inset, d) rate capability of o-SnO₂/MnO₂ and c-SnO₂/MnO₂ nanotube arrays.

electrode structure to enhance the electrochemical performance and rate capability of supercapacitors.

4. Conclusion

In summary, a nano-engineering strategy for fabricating SnO₂/MnO₂ core/shell nanotube arrays is innovated in this work. The fabrication process allows precise controlling of the length, spacing, diameter, wall thickness and the selection of open-end or closed-

end nature of the prepared nanotubes. The synthesized structures exhibit excellent electrochemical performances with an outstanding rate capability, which shows that these well-defined nanostructures are good candidates for positive supercapacitor electrodes in applications that require micro-supercapacitors such as consumer electronics, wireless sensor network, and energy harvesting devices. It should be noted that the presented fabrication process can be applied to many other materials and applications.

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Appendix A. Supplementary data

Supplementary data related to this article can be found at <http://dx.doi.org/10.1016/j.jpowsour.2013.12.029>.

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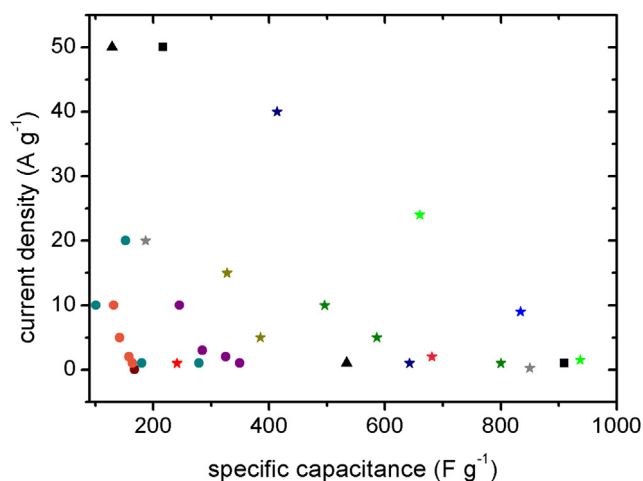


Fig. 5. Values of specific capacitance and current density for various reported MnO₂ based structures (circles: pure MnO₂, stars: core/shell material), including the data of this work: ■ o-SnO₂/MnO₂, ▲ c-SnO₂/MnO₂, ★ MnO₂/Mn [35], ★ MnO₂/TiN [31], ★ MnO₂/TiN [34], ★ MnO₂/SnO₂ [28], ★ MnO₂/Zn₂SnO₄ [33], ★ MnO₂/Au [36], ★ MnO₂/Ni [37], ★ MnO₂/aMEGO [38], ★ MnO₂ NT [39], ★ MnO₂ NW [14], ★ MnO₂ NW [40], ★ CNT/MnO₂ [41].

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